

## 2022-2023 AP Chemistry

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Congratulations on making the decision to take AP Chemistry! This course will move at a fast pace and cover a substantial amount of material, starting with the first day of school. The primary goal of this course is to earn college credit by earning a grade of C or higher for dual enrollment credit or by passing the AP Chemistry exam with a score of 4 or higher in May 2023 (most colleges will not give credit for a score of 1, 2, or 3).

So that we can spend more time on topics new to you in AP Chemistry, you are expected to be familiar answering questions and solving problems using the content covered in your first-year chemistry course. The attached **review assignment** covers first-year chemistry topics that will not be taught in AP chemistry.

Copies of the periodic table and the metric prefixes you will be using in AP Chemistry are included in this assignment. Please note that this periodic table does not include element names. Charges of monatomic ions and key polyatomic ions that need to be memorized by the first test are also included. You are encouraged to make flashcards or use the Quizlet ions card deck (linked on Classroom) to begin learning these ions.

At the end of the first week of class, I will evaluate your completion of this packet and the first assessment in AP Chemistry will follow shortly thereafter (within the first three weeks of class). If you try to review this material, I don't anticipate you having any difficulty with the first test.

In that spirit, enjoy your summer and I look forward to a new year of AP Chemistry with you.

Good luck and have a terrific summer!

Mrs. Amin

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## AP Chemistry Ions

<u>Monatomic Cations</u>	<u>Monatomic Anions</u>	<u>Polyatomic Cations</u>	<u>Polyatomic Anions</u>
<u>Group 1 (including H)</u> H <sup>+1</sup> , hydrogen Li <sup>+1</sup> , lithium Na <sup>+1</sup> , sodium K <sup>+1</sup> , potassium Cs <sup>+1</sup> , cesium  <u>Group 2</u> Be <sup>+2</sup> , beryllium Mg <sup>+2</sup> , magnesium Ca <sup>+2</sup> , calcium Sr <sup>+2</sup> , strontium Ba <sup>+2</sup> , barium  <u>Group 13</u> Al <sup>+3</sup> , aluminum  <u>Transition and Heavier Metals</u> Cr <sup>+2</sup> , chromium (II) Cr <sup>+3</sup> , chromium (III)  Mn <sup>+2</sup> , manganese (II) Mn <sup>+4</sup> , manganese (IV) Mn <sup>+7</sup> , manganese (VII)  Cu <sup>+1</sup> , copper (I) Cu <sup>+2</sup> , copper (II)  Fe <sup>+2</sup> , iron (II) Fe <sup>+3</sup> , iron (III)  Pb <sup>+2</sup> , lead (II) Pb <sup>+4</sup> , lead (IV)  Hg <sup>+2</sup> , mercury (II)  Ni <sup>+2</sup> , nickel (II) Ni <sup>+3</sup> , nickel (III)  Sn <sup>+2</sup> , tin (II) Sn <sup>+4</sup> , tin (IV)  Ag <sup>+1</sup> , silver Zn <sup>+2</sup> , zinc	<u>Group 17 and H</u> H <sup>-1</sup> , hydride F <sup>-1</sup> , fluoride Cl <sup>-1</sup> , chloride Br <sup>-1</sup> , bromide I <sup>-1</sup> , iodide  <u>Group 16</u> O <sup>-2</sup> , oxide S <sup>-2</sup> , sulfide  <u>Group 15</u> N <sup>-3</sup> , nitride P <sup>-3</sup> , phosphide	Ammonium, NH <sub>4</sub> <sup>+1</sup> Mercury (I), Hg <sub>2</sub> <sup>+2</sup>	Acetate, C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-1</sup> Bicarbonate (hydrogen carbonate), HCO <sub>3</sub> <sup>-1</sup> Carbonate, CO <sub>3</sub> <sup>-2</sup>  Perchlorate, ClO <sub>4</sub> <sup>-1</sup> Chlorate, ClO <sub>3</sub> <sup>-1</sup> Chlorite, ClO <sub>2</sub> <sup>-1</sup> Hypochlorite, ClO <sup>-1</sup>  Permanganate, MnO <sub>4</sub> <sup>-1</sup>  Cyanide, CN <sup>-1</sup>  Hydroxide, OH <sup>-1</sup> Peroxide, O <sub>2</sub> <sup>-2</sup>  Nitrate, NO <sub>3</sub> <sup>-1</sup> Nitrite, NO <sub>2</sub> <sup>-1</sup>  Chromate, CrO <sub>4</sub> <sup>-2</sup> Dichromate, Cr <sub>2</sub> O <sub>7</sub> <sup>-2</sup>  Sulfate, SO <sub>4</sub> <sup>-2</sup> Sulfite, SO <sub>3</sub> <sup>-2</sup>  Phosphate, PO <sub>4</sub> <sup>-3</sup> Phosphite, PO <sub>3</sub> <sup>-3</sup>

\*\*\*Note: Transition metals are named with Roman numerals to indicate their oxidation state (charge) if they have multiple oxidation states. Silver and zinc are the only transition metals on this list that have a single oxidation state and therefore are not named with roman numerals. As long as you know which transition metals need Roman numerals, individual charges of these metals do not need to be memorized.

**DO NOT DETACH FROM BOOK.**

## PERIODIC TABLE OF THE ELEMENTS

1 <b>H</b> 1.0079																	2 <b>He</b> 4.0026
3 <b>Li</b> 6.941	4 <b>Be</b> 9.012											5 <b>B</b> 10.811	6 <b>C</b> 12.011	7 <b>N</b> 14.007	8 <b>O</b> 16.00	9 <b>F</b> 19.00	10 <b>Ne</b> 20.179
11 <b>Na</b> 22.99	12 <b>Mg</b> 24.30											13 <b>Al</b> 26.98	14 <b>Si</b> 28.09	15 <b>P</b> 30.974	16 <b>S</b> 32.06	17 <b>Cl</b> 35.453	18 <b>Ar</b> 39.948
19 <b>K</b> 39.10	20 <b>Ca</b> 40.08	21 <b>Sc</b> 44.96	22 <b>Ti</b> 47.90	23 <b>V</b> 50.94	24 <b>Cr</b> 52.00	25 <b>Mn</b> 54.938	26 <b>Fe</b> 55.85	27 <b>Co</b> 58.93	28 <b>Ni</b> 58.69	29 <b>Cu</b> 63.55	30 <b>Zn</b> 65.39	31 <b>Ga</b> 69.72	32 <b>Ge</b> 72.59	33 <b>As</b> 74.92	34 <b>Se</b> 78.96	35 <b>Br</b> 79.90	36 <b>Kr</b> 83.80
37 <b>Rb</b> 85.47	38 <b>Sr</b> 87.62	39 <b>Y</b> 88.91	40 <b>Zr</b> 91.22	41 <b>Nb</b> 92.91	42 <b>Mo</b> 95.94	43 <b>Tc</b> (98)	44 <b>Ru</b> 101.1	45 <b>Rh</b> 102.91	46 <b>Pd</b> 106.42	47 <b>Ag</b> 107.87	48 <b>Cd</b> 112.41	49 <b>In</b> 114.82	50 <b>Sn</b> 118.71	51 <b>Sb</b> 121.75	52 <b>Te</b> 127.60	53 <b>I</b> 126.91	54 <b>Xe</b> 131.29
55 <b>Cs</b> 132.91	56 <b>Ba</b> 137.33	*57 <b>La</b> 138.91	72 <b>Hf</b> 178.49	73 <b>Ta</b> 180.95	74 <b>W</b> 183.85	75 <b>Re</b> 186.21	76 <b>Os</b> 190.2	77 <b>Ir</b> 192.2	78 <b>Pt</b> 195.08	79 <b>Au</b> 196.97	80 <b>Hg</b> 200.59	81 <b>Tl</b> 204.38	82 <b>Pb</b> 207.2	83 <b>Bi</b> 208.98	84 <b>Po</b> (209)	85 <b>At</b> (210)	86 <b>Rn</b> (222)
87 <b>Fr</b> (223)	88 <b>Ra</b> 226.02	†89 <b>Ac</b> 227.03	104 <b>Rf</b> (261)	105 <b>Db</b> (262)	106 <b>Sg</b> (263)	107 <b>Bh</b> (262)	108 <b>Hs</b> (265)	109 <b>Mt</b> (266)	110 <b>§</b> (269)	111 <b>§</b> (272)	112 <b>§</b> (277)	§Not yet named					

*Lanthanide Series	58 <b>Ce</b> 140.12	59 <b>Pr</b> 140.91	60 <b>Nd</b> 144.24	61 <b>Pm</b> (145)	62 <b>Sm</b> 150.4	63 <b>Eu</b> 151.97	64 <b>Gd</b> 157.25	65 <b>Tb</b> 158.93	66 <b>Dy</b> 162.50	67 <b>Ho</b> 164.93	68 <b>Er</b> 167.26	69 <b>Tm</b> 168.93	70 <b>Yb</b> 173.04	71 <b>Lu</b> 174.97
†Actinide Series	90 <b>Th</b> 232.04	91 <b>Pa</b> 231.04	92 <b>U</b> 238.03	93 <b>Np</b> 237.05	94 <b>Pu</b> (244)	95 <b>Am</b> (243)	96 <b>Cm</b> (247)	97 <b>Bk</b> (247)	98 <b>Cf</b> (251)	99 <b>Es</b> (252)	100 <b>Fm</b> (257)	101 <b>Md</b> (258)	102 <b>No</b> (259)	103 <b>Lr</b> (260)

INFORMATION IN THE TABLE BELOW AND IN THE TABLES ON PAGES 3-5 MAY BE USEFUL IN ANSWERING THE QUESTIONS IN THIS SECTION OF THE EXAMINATION.

# Metric Conversions

Unit	Symbol	*Equivalent Expressions*	
mega	M	1 Mg = 1,000,000 g = $10^6$ g	1 Mg = 1,000,000 g = $10^6$ g
kilo	k	1 kg = 1,000 g = $10^3$ g	1 kg = 1,000 g = $10^3$ g
hecta	h	1 hg = 100 g = $10^2$ g	1 hg = 100 g = $10^2$ g
deca	da	1 dag = 10 g = $10^1$ g	1 dag = 10 g = $10^1$ g
o		1g = $10^0$ g	1g = $10^0$ g
deci	d	1 g = 10 dg = $10^1$ dg	1 dg = 0.1 g = $10^{-1}$ g
centi	c	1 g = 100 cg = $10^2$ cg	1 cg = 0.01 g = $10^{-2}$ g
milli	m	1 g = 1,000 mg = $10^3$ mg	1 mg = 0.001 g = $10^{-3}$ g
micro	$\mu$	1 g = 1,000,000 $\mu$ g = $10^6$ $\mu$ g	1 $\mu$ g = 0.000001 g = $10^{-6}$ g
nano	n	1 g = 1,000,000,000 ng = $10^9$ ng	1 ng = 0.000000001 g = $10^{-9}$ g
pico	p	1 g = 1,000,000,000,000 pg = $10^{12}$ pg	1 pg = 0.000000000001 g = $10^{-12}$ g

\* Any quantity can be substituted for g; ie. 1 L = 1000 mL just as 1 g = 1000 mg

**A helpful pnemonic for memorizing prefixes (you need to know these):**

**Many kids have dropped over dead converting metric measurements in problems.**

## Advanced Placement Chemistry Review Assignment

### Topic 1: Significant Figures & Scientific Notation

- Count the number of significant figures in the following measurements.
  - 2.71 g \_\_\_\_\_
  - 0.00047 kg \_\_\_\_\_
  - $7.0 \times 10^5$  m \_\_\_\_\_
  - 1,030 L \_\_\_\_\_
  - 150 pencils \_\_\_\_\_
  - 37500  $\mu\text{g}$  \_\_\_\_\_
  - 0.1010 cm \_\_\_\_\_
- Express each of the following in proper scientific notation (Pay attention to sig figs and units).
  - 0.000125 m \_\_\_\_\_
  - 155.0 mL \_\_\_\_\_
  - 123,030,000 ng \_\_\_\_\_
  - $481.9 \times 10^{-9}$  cm \_\_\_\_\_
- Calculate the correct answer with proper units and significant figures for each of the following:
  - $12 \text{ g} + 0.677 \text{ g} + 86.33 \text{ g} =$  \_\_\_\_\_
  - $(355.78 \text{ g}) / (0.056 \text{ g}) =$  \_\_\_\_\_
  - $97.34 \text{ mL} - 34.1 \text{ mL} =$  \_\_\_\_\_
  - $14.68 \times 5 =$  \_\_\_\_\_
- Perform the following calculations with scientific notation and report your answer with the correct number of significant figures.
  - $0.14 \times (6.02 \times 10^{23}) =$  \_\_\_\_\_
  - $\frac{(9.875 \times 10^4) - (9.795 \times 10^4)}{9.875 \times 10^4} \times 100 \% =$  \_\_\_\_\_ (assume 100 is exact)
  - $\frac{(3.8 \times 10^{-12}) \times (4.0 \times 10^{-13})}{(4 \times 10^{12}) \times (6.3 \times 10^{13})} =$  \_\_\_\_\_

### Topic 2: Dimensional Analysis

Show work using dimensional analysis. No work = no credit even if answer is correct. Follow significant figures and rounding rules unless the number of significant figures is specified. Include units where appropriate.

- How many hours are in a week? Report your answer to three significant figures.
- Find the number of centimeters in  $1.00 \times 10^2$  yards. (1 yd = 3 ft, 1 ft = 12 in, 2.54 cm = 1 in)
- If Jules Verne expressed the title of his famous book, Twenty Thousand Leagues under the Sea in basic SI units, what would the title be? Round your answer to three significant figures. (1 league = 3.45 mi, 1 mi = 1609 m)
- How many  $\mu\text{L}$  are present in 250 mL of  $\text{H}_2\text{O}$ ?

9. Wavelengths are often represented in nm. What is the diameter of a helium (He) atom in nm if it is equivalent to  $1.0 \times 10^{-13}$  km?
10. The area of a rectangular room has a length of 10.5 m and a width of 4.50 m. What is this area in  $\text{m}^2$ ? In  $\text{cm}^2$ ?
11. The acceleration of a sphere is determined to be  $9.52 \text{ m/s}^2$ . What is the acceleration in  $\text{km/min}^2$ ?

### **Topic 3: Density and Temperature**

Show all work. No work = no credit even if answer is correct. Follow significant figures and rounding rules. Include units where appropriate.

12. A rectangular block has dimensions of 2.9 cm x 3.5 cm x 10.0 cm. The mass of the block is 615.0 grams. What are the volume and the density of the block?
13. The density of pure silver is  $10.5 \text{ g/mL}$  at  $20^\circ\text{C}$ . If 5.25 grams of pure silver pellets are added to a graduated cylinder containing 11.2 mL of water, to what volume will the water in the cylinder rise?
14. You can figure out whether a substance floats or sinks if you know its density and the density of the liquid. In which of the liquids listed below will high-density polyethylene, HDPE, float? HDPE, a common plastic, has a density of  $0.97 \text{ g/cm}^3$ . It does not dissolve in any of the following

liquids.

Substance	Density (g/cm <sup>3</sup> )
ethylene glycol	1.1088
water	0.9997
ethanol	0.7893
methanol	0.7914
acetic acid	1.0492
glycerol	1.2613

15. Mercury is found as a liquid at room temperature. If it has a boiling point of 630. K, what is this boiling point in degrees Celsius?

#### **Topic 4: Precision and Accuracy**

16. The density of ethanol was determined experimentally at 25°C in a series of trials to be 0.608 g/mL, 0.705 g/mL, and 0.689 g/mL. The accepted density of ethanol is reported to be 0.789 g/mL.

- a. Are the experimental densities precise? Why/Why not?
  
  
  
  
  
  
  
  
  
  
- b. Calculate % error for this experiment. Use the average experimental density in your calculation and report your answer to 0.1%. Show your work.
  
  
  
  
  
  
  
  
  
  
- c. Are the experimental densities accurate? Why/Why not?

#### **Topic 5: Properties and Changes**

17. Categorize each of the following as an element, a compound, or a mixture:

- a. carbonated water \_\_\_\_\_
- b. tungsten \_\_\_\_\_
- c. aspirin (acetylsalicylic acid) \_\_\_\_\_
- d. air \_\_\_\_\_
- e. lye (sodium hydroxide) \_\_\_\_\_
- f. fluorine \_\_\_\_\_

18. Iron pyrite, also known as fool's gold, has a shiny golden metallic appearance. Crystals are often in the form of perfect cubes. A cube of iron pyrite measuring 0.40 cm on each side has a mass of 0.064 g.

- a. Which of these observations are qualitative and which are quantitative?

- b. Which of these observations are extensive (dependent on the amount of substance present) and which are intensive (independent of the amount of substance present)?
19. Identify the following as a physical property, physical change, chemical property, or chemical change:
- Ethanol has a density of 0.697 g/mL. \_\_\_\_\_
  - The solution turns blue upon mixing water and food coloring. \_\_\_\_\_
  - Wood burns in an oven. \_\_\_\_\_
  - Methyl alcohol is highly flammable. \_\_\_\_\_
  - Ice melts in a beaker. \_\_\_\_\_
  - Methyl ethanoate smells like apples. \_\_\_\_\_
  - Iron rusts on a car. \_\_\_\_\_
  - Alkali metals react strongly in hydrochloric acid. \_\_\_\_\_

### **Topic 6: Atom Structure & History**

20. How many protons and neutrons are contained in the nucleus of each of the following atoms? How many electrons are present in each of these neutral atoms?

- $^{13}_6\text{C}$       \_\_\_\_ protons      \_\_\_\_ neutrons      \_\_\_\_ electrons
- $^{208}_{82}\text{Pb}$       \_\_\_\_ protons      \_\_\_\_ neutrons      \_\_\_\_ electrons

21. Complete the following table:

Name	Mass #	Atomic #	# of Protons	# of Neutrons	# of Electrons	Symbol
Gallium-70					31	
						$^{31}_{15}\text{P}^{-3}$
Strontium-80					36	
						$^{55}_{25}\text{Mn}^{+2}$

22. The natural abundance for boron isotopes is 19.9% boron-10 (exact mass 10.013 amu) and 80.1% boron-11 (exact mass 11.009 amu). Calculate the average atomic mass of boron using the exact masses instead of mass numbers in your calculations. Show your work. Follow significant figures and rounding rules. Include appropriate units.

23. Europium has two stable isotopes,  $^{151}\text{Eu}$  and  $^{153}\text{Eu}$ , with masses of 150.9197 u and 152.9212 u, respectively. Calculate the percent abundances of these isotopes of europium to 0.1%. Hint: The percent abundances of these two isotopes must add to 100%. Show your work. Follow significant figures and rounding rules. Include appropriate units.

24. Identify the scientist(s) noted for the following events in atomic history.
- identified the electron; noted for the plum pudding model \_\_\_\_\_
  - noted for the first atomic theory of the atom; solid sphere model \_\_\_\_\_
  - developed the planetary model; electrons in fixed orbits \_\_\_\_\_
  - developed the quantum mechanical model; electrons are localized to orbitals \_\_\_\_\_
  - identified the proton and the nucleus; nuclear model \_\_\_\_\_
  - determined the charge of an electron \_\_\_\_\_
  - described wave theory \_\_\_\_\_
  - known for the uncertainty principle \_\_\_\_\_
  - developed quantum numbers \_\_\_\_\_

25. Identify the model of the atom described in the following statements.

- currently accepted model \_\_\_\_\_
- model that first included a subatomic particle \_\_\_\_\_
- model developed using the gold foil experiment \_\_\_\_\_
- original model of the atom; atom was thought to be "indivisible" \_\_\_\_\_
- model that only showed the movement of hydrogen's electron accurately; involved "quants"  
\_\_\_\_\_

### **Topic 7: Periodic Table Structure**

Identify by name the group or section of the periodic table noted for the following features.

26. a. group containing the most reactive nonmetals; all are diatomics; form -1 ions  
\_\_\_\_\_

b. group containing metals that only form +2 ions \_\_\_\_\_

c. set of metals that often form colored ions in solution; the majority have multiple charges as ions  
\_\_\_\_\_

d. group containing the most reactive metals; form +1 ions \_\_\_\_\_

e. group containing least reactive elements on periodic table, typically inert  
\_\_\_\_\_

27. These elements start with the letter B: B, Ba, Bk, Bi, and Br. Identify which of these elements match the following descriptions. You may use elements once, more than once, or not at all.

- Which are metals? \_\_\_\_\_
- Which are liquids? \_\_\_\_\_
- Which are actinides? \_\_\_\_\_
- Which are main block elements? \_\_\_\_\_

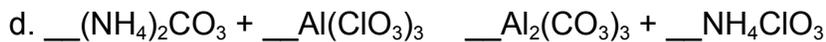
### **Topic 8: Compound Nomenclature**

28. Name or give the formula for the following compounds. All ions included in the summer letter are required to be memorized by name and by formula.

<u>Name</u>	<u>Formula</u>
a. lithium fluoride	_____
b. _____	K <sub>2</sub> O
c. calcium phosphate	_____
d. _____	MnCl <sub>2</sub>
e. silver sulfide	_____
f. _____	Cu <sub>2</sub> O
g. aluminum sulfate	_____
h. _____	ZnCO <sub>3</sub>
i. chromium (III) phosphide	_____
j. _____	SO <sub>3</sub>
k. lead (IV) hydroxide	_____
l. _____	N <sub>2</sub> O <sub>5</sub>
m. ammonium sulfite	_____
n. _____	BaCr <sub>2</sub> O <sub>7</sub>
o. sodium peroxide	_____
p. _____	NH <sub>3</sub> (use common names; see ppt/videos if necessary)
q. nickel (II) hypochlorite	_____
r. _____	Fe(CN) <sub>3</sub>
s. rubidium chromate	_____
t. _____	Mg <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>

### **Topic 9: Equations**

29. Balance the following equations using the lowest whole-number coefficients.

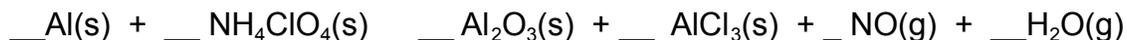


30. Write balanced chemical equations for the following word equations. Use the lowest possible whole-number coefficients to balance the equations.
- Aqueous solutions of ammonium sulfate and barium nitrate form a precipitate of barium sulfate and aqueous ammonium nitrate.
  - Elemental magnesium and oxygen gas combine to form solid magnesium oxide.
  - Chlorine gas and aqueous potassium bromide react to form bromine liquid and aqueous potassium chloride.
  - Solid copper (II) carbonate decomposes to form crystals of copper (II) oxide and carbon dioxide gas.
  - Sulfuric acid is neutralized by lithium hydroxide to form water and aqueous lithium sulfate.
  - Liquid benzene,  $C_6H_6$ , undergoes combustion in oxygen gas, making carbon dioxide gas and steam.

**Topic 10: Mole Conversions & Stoichiometry**

Show your work. No work = no credit. Follow significant figures and rounding rules. Include appropriate units.

31. a. Calculate the number of moles in 500. atoms of iron (Fe).
- b. What is the molar mass of lead (IV) carbonate,  $Pb(CO_3)_2$ ?
- c. How many formula units are present in 87.2 grams of lead (IV) carbonate?
- d. What percentage of oxygen is found in lead (IV) carbonate? Round your answer to 0.1%.
32. The reusable booster rockets of the U.S. space shuttle employed a mixture of aluminum and ammonium perchlorate for fuel. A possible reaction for this is:



- a. Balance the above reaction using the lowest possible whole-number coefficients.
  - b. If 4.00 g of aluminum reacted completely, how many grams of aluminum oxide would be made?
  
  - c. If 4.18 g of aluminum chloride was produced, how many moles of ammonium perchlorate would be consumed?
  
  - d. How many molecules of nitrogen monoxide would form if  $6.3 \times 10^{25}$  formula units of aluminum oxide were also produced?
33. The decomposition of ammonia is shown in the following equation:  $2\text{NH}_3\text{(g)} \rightarrow \text{N}_2\text{(g)} + 3\text{H}_2\text{(g)}$ .
- a. 42.0 g of nitrogen has what volume in liters at STP?
  
  - b. 150 L of  $\text{NH}_3$  undergoes decomposition to form how many liters of hydrogen gas at STP?
  
  - c. How many liters of ammonia were decomposed at STP if  $3.0 \times 10^{23}$  nitrogen molecules were made?

In the space below, write the electron configurations of the following elements:

1. Oxygen \_\_\_\_\_

2. Chlorine \_\_\_\_\_

3. Argon \_\_\_\_\_

4. Iron \_\_\_\_\_

Determine what elements are denoted by the following electron configurations:

11)  $1s^2 2s^2 2p^6 3s^2 3p^5$  \_\_\_\_\_

12)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^1$  \_\_\_\_\_

13)  $1s^2 2s^2 2p^3$  \_\_\_\_\_

Draw the orbital diagrams for the following elements: Example: Mg (12 e<sup>-</sup>)

\_\_\_\_\_

1s    2s    2p                    3s

16) Nitrogen

17) Sodium

18) Chlorine

19) Potassium

20) Iron

## Topic 12: Periodic Trends

On the blank periodic table, color and label:

- a. alkali metals
- b. alkaline metals
- c. transition metals
- d. nonmetals
- e. metalloids
- f. halogens
- g. noble gases
- h. inner transition metals

On the blank periodic table, color and label.

- i. the s block
- j. the p block
- k. the d block
- f. the f block

2. On the blank periodic table, draw arrows to show the following periodic trends across each period and down each group. Be sure to label which way the trend is increasing and which way it is decreasing.

- a. Atomic radius
- b. Ionization energy
- c. Electronegativity

# Periodic Table of the Elements

	2											13	14	15	16	17	
		3	4	5	6	7	8	9	10	11	12						
		*															
		**															

3

4

5

6


\*

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**Directions: Use your notes to answer the following questions.**

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.
2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.
3. Why does fluorine have a higher ionization energy than iodine?
4. Why do elements in the same family generally have similar properties?
5. Indicate whether the following properties increase or decrease from left to right across the periodic table.
  - a. atomic radius (excluding noble gases)
  - b. first ionization energy
  - c. electronegativity
6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?
7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?
8. Circle the atom in each pair that has the largest atomic radius.

a. Al or B	c. Na or Al	e. S or O
b. O or F	d. Br or Cl	f. Mg or Ca
9. Circle the atom in each pair that has the greater ionization energy.

a. Li or Be	c. Ca or Ba	e. Na or K
b. P or Ar	d. Cl or Si	f. Li or K
10. Circle the atom in each pair that has the greater electronegativity.

a. Ca or Ga	c. Br or As	e. Li or O
b. Ba or Sr	d. Cl or S	c. O or S

